

PART – I [Attempt all questions from this part – 19 Marks]

Q.No.I Select the correct answers to the questions from the given options. (16)

- i. When was the sitting of the Constituent Assembly held?
 - a. December 5, 1946
 - b. December 9, 1946
 - c. December 25, 1945
 - d. December 24, 1946
- ii. Who proposed the 'Objectives Resolution'?
 - a. Dr Rajendra Prasad
 - b. Subhas Chandra Bose
 - c. Mahatma Gandhi
 - d. Jawahar Lal Nehru
- iii. On the basis of which plan was the Constituent Assembly constituted?
 - a. Cabinet Mission Plan
 - b. Cripps Mission
 - c. Mountbatten Plan
 - d. None of the above
- iv. Bronze statues were made by a special process called _____.
 - a. moulded clay process
 - b. clay wax process
 - c. lost wax process
 - d. lost clay process
- v. Which of the following is a feature of the Great Bath?
 - a. Tank in a courtyard
 - b. Porticos and rooms
 - c. Tank surrounded by corridors
 - d. All of the above
- vi. 'The Dancing Girl' has which of the following features?
 - a. Stone sculpture
 - b. half-closed eyes
 - c. heavily-bangled left arm
 - d. seated cross-legged
- vii. Seals are not a source of information for
 - a. Script
 - b. trade
 - c. beliefs
 - d. society
- viii. Seals were made of
 - a. Steatite
 - b. Terracotta
 - c. Agate
 - d. All of the above
- ix. Read the statements given below and select the option that shows the correct relationship between [A] and [R].

Assertion [A] : The constitution of India was adopted and passed by the Constituent Assembly on November 26, 1949.

Reason [R] : After signing of the Constitution all the clauses came into force with immediate effect.

 - a. [A] is true but [R] is false.
 - b. [A] is false but [R] is true.
 - c. Both [A] and [R] are true but [R] is not the correct explanation of [A].
 - d. Both [A] and [R] are false.

- Q.No.II
- i. What is meant by the term Constitution? (2)
 - ii. Which clauses came into force after signing of the Constitution? (2)
 - iii. Name the two important sources of information about the Harappan Civilization. (2)
 - iv. Name any two causes that led to the decline of the Harappan Civilization. (2)
 - v. Name the four animals depicted on the Pashupati seal. (2)

[P.T.O.]

CARMEL SCHOOL , MADHUPUR
FIRST UNIT TEST 2025 – 2026

STD: X

Time: 1 hr

CHEMISTRY

M.M: 50

Answer the paper be written on the paper provided separately.

Section – A is compulsory. Attempt all questions from Section – B.
The intended marks for questions or part of questions are given in bracket [].

Section – A [Attempt all the questions from this section.]

Q.No.I Choose the correct answer to the question from the given options. [10]

- I. An element in period 3, whose electron affinity is zero:
a. Neon b. Sulphur c. Sodium d. Argon
- II. Element with an atomic number 19 will:
a. Accept an electron and get oxidized c. Accept an electron and get reduced
b. lose an electron and get oxidized d. lose an electron and get reduced
- III. Most favourable condition for electrovalent bonding are:
a. Low ionization energy of one atom and high electron affinity of the other element
b. High electron affinity and high ionization energy of both the elements
c. low electron affinity and low ionization energy of both the elements
d. high ionization energy of one atom and low electron affinity of the other element
- IV. How many covalent bond, coordinate bonds and lone pair of electrons are present in hydronium ion:
a. 1 covalent bond, one co-ordinate bond, one lone pair
b. 1 covalent bond, two co-ordinate bond, two lone pair
c. 2 covalent bond, one co-ordinate bond, one lone pair
d. 2 covalent bond, 2 co-ordinate bond, one lone pair
- V. Lead chloride is a white ppt. which is:
a. Soluble in hot water but insoluble in cold water
b. Insoluble in hot water but soluble in cold water
c. Soluble in hot and cold water
d. Insoluble in hot water and cold water
- VI. the metallic oxide which when reacts with HCl forms salt and water:
a. carbon monoxide b. nitrous oxide
b. ammonium hydroxide c. sodium oxide
- VII. an aqueous solution of ammonia is:
a. Neutral b. acidic c. basic d. amphoteric
- VIII. Promoter used in Haber's process:
a. Mo b. Fe c. Pt d. Ni
- IX. A compound which on reaction with HCl leaves yellow residue behind:
a. AgCl_2 b. Conc. H_2SO_4 c. MnO_2 d. $\text{Na}_2\text{S}_2\text{O}_3$
- X. The color of the precipitate obtained on adding ammonium hydroxide in a small quantity of iron(III) chloride is :
a. White b. Dirty green c. Pale blue d. Reddish brown

Q.No.II Name the following: [5]

- a. The gas produced when excess ammonia reacts with chlorine.
- b. Two compounds of lead which react with concentrated hydrochloric acid to liberate chlorine.
- c. An ion formed by the loss of electron from neutral atom.

[P.T.O.]



- d. Two gases which gives dense white fumes with ammonia.
- e. Name the first and last element in period 3.

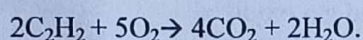
Q.No.III Match the following column A with Column B.

[5]

Column – A	Column – B
a. Electron affinity	1. Tendency to loose electron
b. Ionisation energy	2. Verticle column
c. Metallic character	3. Energy released
d. Non metallic character	4. Horizontal rows
e. Group	5. Tendency to gain electron
	6. Energy required

Section – B [Attempt all questions]

Q.No.IV I. Oxygen oxidizes Ethyne to carbon dioxide and water as shown by the equation:



What volume of Ethyne gas at STP is required to produce 8.4 dm^3 of carbondioxide at STP. [2]

2. Identify the gas evolved and give the chemical test for each of the following cases: [2]
 - a. Dil. HCl reacts with sodium sulphite.
 - b. Dil. HCl reacts with Iron II sulphide.
3. Fill in the blanks from the choice given within the brackets: [3]
 - a. Ammonia is burnt in atmosphere of oxygen with _____ (green/ bluish white) color.
 - b. Ammonium salts react with _____ (acid/ alkalis) to liberate ammonia.
 - c. If an element has low ionization energy than it is likely to be ____ (metallic/non metallic).
4. the following question are related with certain property of an element X having atomic number 20: [3]
 - a. Identify element X?
 - b. Which group and period element X is present?
 - c. Is this element greater in size than magnesium greater or smaller?

Q.No.V

I. The table below shows electronic arrangement of six atoms A to F :

[2]

Atom	A	B	C	D	E	F
Electronic configuration	2,5	2	2, 6	2,8, 6	2,8,8	2,8,3

With respect to the table select the following:-

[3]

- a. Two atoms from same group of the periodic table?
- b. The atom which is most electronegative?
- c. Two nobel gases?

II. The metal of group 2 from top to bottom are Be, Mg, Ca, Sr and Ba.

[2]

- (a) Which of these elements will form ions most readily and why?
- (b) State one common feature in electronic configuration of all these elements?

III. When a compound X a black colour substance reacts with conc. Hydrochloric acid a greenish yellow gas Y is released. Name the substance X and Y and also write a balanced chemical equation for the same.

[2]

[P.T.O]

IV. Answer the following questions:-

[3]

- Draw the electrton dot structure of ammonium ion.
- Name the bond formed by shared pair of electrons with both the electrons coming from the same atoms?
- Among the following compounds , identify the compounds that has all three bonds(ionic, covalent coordinate bond)

(i) ammonia	(ii) ammonium chloride
(iii) sodium hydroxide	(iv) calcium chloride

Q.No.VI 1. (a) If 6 litres of hydrogen and 4 litres of chlorine are mixed and exploded and if the water is added to the gases formed. Find the volume of residual gas? [2]

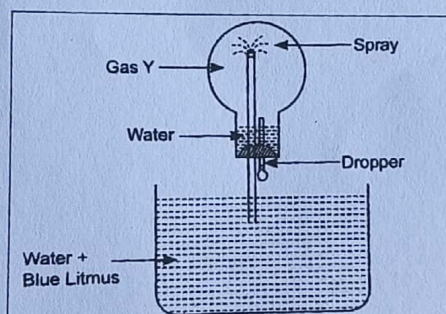
(b) State avagadro's law? [2]

2. In the laboratory preparation of Hydrochloric acid HCl gas is dissolved in water. [3]

- Draw a diagram to show the arrangement used for the absorption of HCl in water.
- Why is such an arrangement necessary. Give two reasons?
- Write the chemical equation for the laboratory preparation of HCl gas when the reactant on heated

(i) below 200°C	(ii) above 200°C .
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3. Study the figure given below and answer the question that follows:



[3]

- Identify the gas Y?
- What property a gas Y does this experiment demonstrate?
- Name another gas which has the same property and can be demonstrated through same experiment?

□□□□□□□□

